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(ii) **Nitrogen molecule / Molekul nitrogen:**

(a) Nitrogen atom with an electron arrangement 2.5 needs 3 electrons to achieve stable octet electron arrangement. / Atom nitrogen dengan susunan elektron 2.5 memerlukan 3 elektron untuk mencapai susunan elektron oktet yang stabil.

(b) Two nitrogen atoms share 3 pairs of electrons to achieve a stable octet electron arrangement, form a nitrogen molecule. Each nitrogen atom achieves stable octet electron arrangement.

Dua atom nitrogen berkongsi 3 pasang elektron untuk mencapai susunan elektron oktet yang stabil, membentuk satu molekul nitrogen. Setiap atom nitrogen mencapai susunan elektron oktet yang stabil.

Draw the electron arrangement of the molecule formed. / Lukiskan susunan elektron bagi molekul yang terbentuk.



The number of electron pairs shared is 3 pairs. **Triple covalent bond** is formed.

Bilangan pasangan elektron dikongsi adalah 3 pasang. Ikatan kovalen **ganda tiga** terbentuk.



Steps to Explain the Formation of Covalent Bond for Covalent Compound

Langkah untuk Menerangkan Pembentukan Ikatan Kovalen dalam Sebatian Kovalen

Predict formula / Meramal formula

Determine the formula of the covalent compound formed.

Tentukan formula sebatian kovalen yang terbentuk.

Draw the electron arrangement for the compound formed based on the formula. / Lukis susunan elektron untuk sebatian yang terbentuk berdasarkan formula.

Lukiskan susunan elektron untuk sebatian yang terbentuk berdasarkan formula.

- Write the electron arrangement of each atom.
Tulis susunan elektron untuk setiap atom.
- By referring to the number of valence electrons for each atom, determine the number of electrons needed by each atom to achieve stable duplet/octet electron arrangement.
Dengan merujuk kepada bilangan elektron valens setiap atom, tentukan bilangan elektron yang diperlukan oleh setiap atom untuk mencapai susunan elektron oktet/duplet yang stabil.
- Cross the number of electrons needed to achieve stable duplet/octet electron arrangement. / Silangkan bilangan elektron yang diperlukan untuk mencapai susunan elektron duplet/octet yang stabil.

- Electrons are shared in pair.
Elektron dikongsi dalam pasangan.
- The number of pairs of electrons shared is based on the number of electrons needed by each atom to achieve stable duplet/octet electron arrangement.
Bilangan pasangan elektron yang dikongsi adalah berdasarkan bilangan elektron yang diperlukan oleh setiap atom untuk mencapai susunan duplet/octet yang stabil.

Explanation / Penjelasan

(c) State the electron arrangement achieved by each atom in the molecule (stable duplet/octet electron arrangement).

Nyatakan susunan elektron yang dicapai oleh setiap atom dalam molekul (susunan elektron duplet/octet yang stabil).

(b) Based on the drawing of the electron arrangement, state:

- Berdasarkan lukisan susunan elektron, nyatakan:*
- * the number of pairs of electrons shared between atoms. / bilangan pasangan elektron yang perlu dikongsi antara atom.
 - * the number of electrons of each atom contribute for sharing in the molecule and type of covalent bond formed (single/double). bilangan elektron yang disumbangkan oleh setiap atom untuk dikongsi oleh setiap atom dalam molekul dan jenis ikatan kovalen yang terbentuk (tunggal/ganda dua).

(a) State the electron arrangement of each atom and the number of electrons needed to achieve stable duplet/octet electron arrangement.

Nyatakan susunan elektron setiap atom dan bilangan elektron yang diperlukan oleh setiap atom untuk mencapai susunan elektron duplet/octet yang stabil.



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